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Review Answers

Objectives: Given a reaction (described in words), be able to start with at any of the three starting points on the flow chart (mass A, volume A(aq), volume A(g)) and calculate any of the four possible outcomes of the flow chart (mass B, volume B(aq), volume B(g),

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molarity B).

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1. Calculation of
quantities in

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Answers 1. a. Na_3PO_4
b. $\text{Ca}(\text{NO}_3)_2$ Na = 3
mol x 22.99 g/mol =
68.97 g Ca = 1 mol x
40.08 g/mol = 40.08 g
P = 1 mol x 30.97
g/mol = 30.97 g N = 2
mol x 14.01 g/mol =
28.02 g O = 4 mol x
16.00 g/mol = 64.00 g
O = 6 mol x 16.00
g/mol = 96.00 g
163.94 g 164.10 g ...

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Answer Key Example 1:
Calculate the mass of a
magnesium, Mg, atoms
in grams. 24.035 g
Mg.1 mol Mg.1
molecule Mg = 4.04 x
10⁻²³ g/Mg atom 1 mol
Mg 6.02 x
10²³ molecules 1 atom
Mg

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CHAPTER 9 REVIEW
Stoichiometry SECTION
3 PROBLEMS Write the
answer on the line to
the left Show all your
work in the space
provided 1 88% The
actual yield of a
reaction is

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CHAPTER 9 REVIEW

Stoichiometry SECTION

3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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class): *Note: Answer
to 7a should be 1.45×10^{25} atoms

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The study of the
quantitative

relationships between
the amounts of

reactants used and the
amounts of products

formed by a chemical
reaction is called

stoichiometry. ____ 2.

Stoichiometry is based
on the law of

conservation of mass.

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3 PROBLEMS Write the
answer on the line to
the left. Show all your
work in the space
provided. 1. _____ The
actual yield of a
reaction is 22 g and
the theoretical yield is

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