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Scientific Buffer:

Buffers are defined as solutions which resists small change in pH by adding small amount of acid or base. A buffer usually consists of a weak acid and its salt

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(fore eg, acetic acid  
and sodium acetate) ...

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Scientific Set the  
100-mL beaker  
containing the acetic  
acid-acetate buffer  
solution on a magnetic  
stirrer, if one is  
available. Add a stir  
bar to the solution.  
Gently stir the buffer

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solution. 7. Place the pH electrode in the solution.

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with Base Set Up:  
Calculations: Graph  
and table NaOH (the  
pH was after the pH

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meter was removed)

Add 55 mL of 0.5 M acetic acid with 45 mL of 0.5 sodium acetate to create 100 mL buffer solution.

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its buffer capacity.

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solution (more  
precisely, pH buffer or  
hydrogen ion buffer) is

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an aqueous solution consisting of a mixture of a weak acid and its conjugate base, or vice versa. Its pH changes very little when a small amount of strong acid or base is added to it.

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Properties of Buffer Solution. Buffer solutions are certainly resistant to changes in pH. However, the pH of a buffer solution can change if there is an addition of sufficient strong acid or strong base. Buffer capacity refers to the amount of strong acid or base a buffer solution can take before significant pH changes take place.

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**What is Buffer**

**Solution? -**

**Definition,  
Application,  
Properties**

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solution (more  
precisely, pH buffer or  
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an aqueous solution  
consisting of a mixture  
of a weak acid and its  
conjugate base, or vice

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solutions. The students  
make two ideal buffer  
solutions, one  
consisting of a weak  
acid and its conjugate

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pH Properties of Buffer Solutions. Flinn AP Chemistry Laboratory #19. Introduction. One of the most important applications of acids and bases in chemistry and biology is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it.

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Ph Properties Of Buffer Solutions A buffer solution (more precisely, pH buffer or hydrogen ion buffer) is an aqueous solution consisting of a mixture of a weak acid and its conjugate base, or vice versa. Its pH changes very little when a small amount of strong acid or base is added to it. Buffer solutions are used as a means of

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