

# Chapter 5 Electrons In Atoms Guided Reading Answers

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Chapter 5 Electrons in Atoms. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. SmileyKylie0923. Key Concepts: Terms in this set (57) Dalton. The atom is a tiny, indestructible particle with no internal structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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116 Chapter 5 Electrons in Atoms CHAPTER 5 What You'll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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## **Chapter 5 - Electrons in Atoms**

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how: Aufbau principle - electrons enter the lowest energy first. This causes difficulties because of the overlap of orbitals of different energies - follow the diagram! Pauli Exclusion Principle - at most 2 electrons per orbital ...

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Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets

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move around the sun. Rutherford's model fails to explain why objects change color when heated.

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136 Chapter 5 • Electrons in Atoms Section 5.1.1 Figure 5.1 Different elements can have similar reactions with water. Objectives Compare the wave and particle natures of light. Define a quantum of energy, and explain how it is related to an energy change of matter. Contrast continuous electromagnetic

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their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of  $4.90 \times 10^{-7} \text{ m}$ . What is the frequency of the light? ...

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Chapter 5 "Electrons in Atoms ... Describe the energies and positions of electrons according to the quantum mechanical model. Section 5.1 Models of the Atom. OBJECTIVES: Describe how the shapes of orbitals related to different sublevels differ. Ernest Rutherford's Model.

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